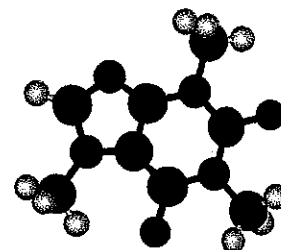


## Assignment Sheet 3 – CP Chemistry Modern Chemistry Chapters 5 + 6



You are expected to READ each chapter in the textbook. Then complete the assigned questions for each chapter. The chapter homework is due approximately once per week.

When doing your homework, be sure to answer each question completely. Any question asking for an explanation should be answered in a complete sentence. Any mathematical question must show work.

While I will always expect you to know the vocabulary from the chapters, I will NOT be counting this as part of your homework grade. However, be sure to study the vocabulary for the tests and quizzes.

Due Date

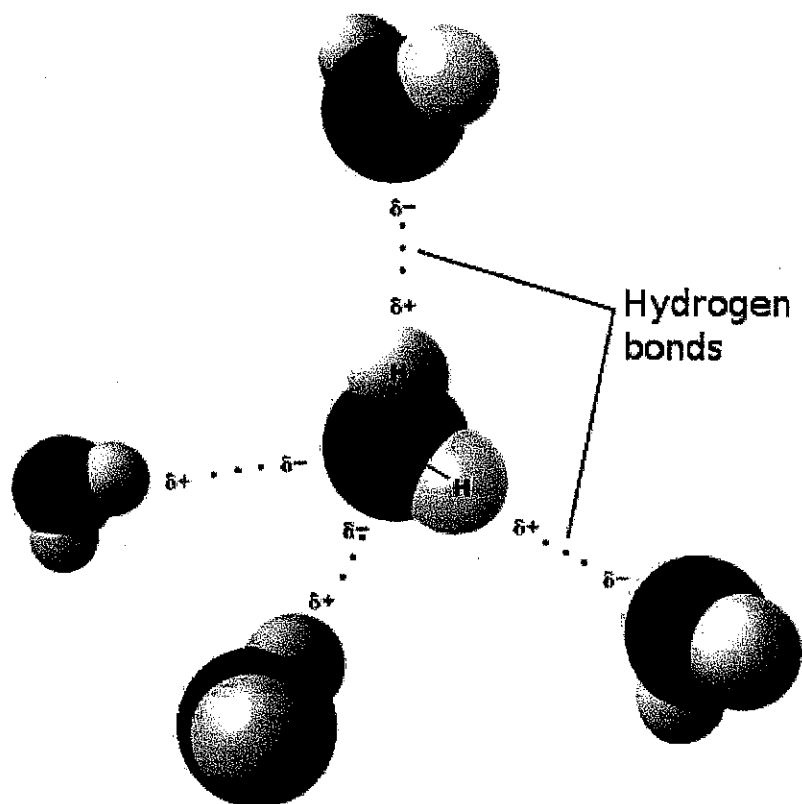
### Chapter 5: Atoms: The Periodic Law

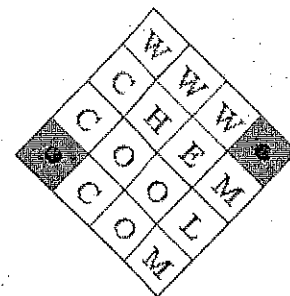
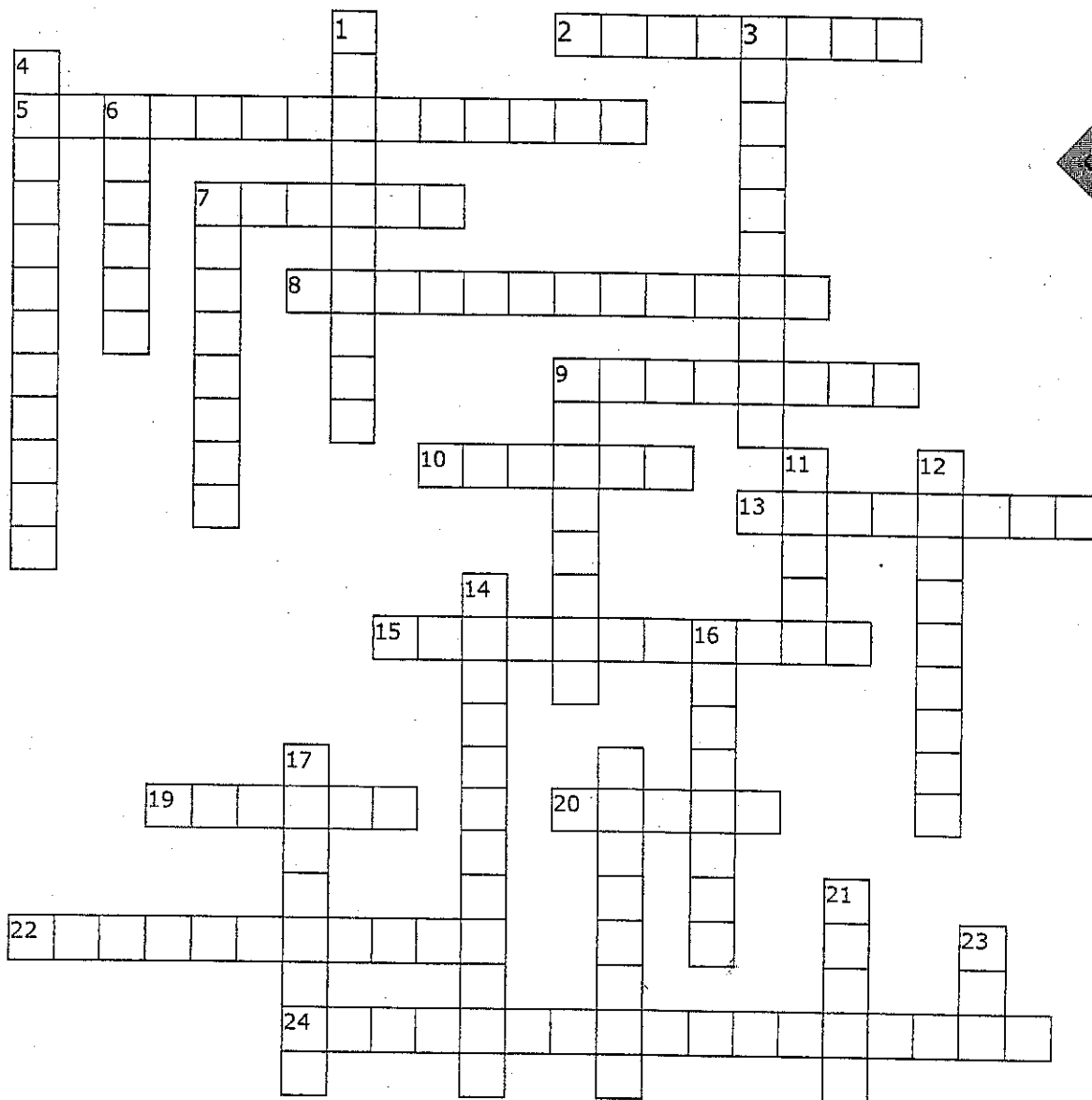
Page 166: #1, 2, 4, 6-9, 13-15, 19, 22-25, 27-31, 33, 35-38, 43, 45

Due Date

### Chapter 6: Chemical Bonding

Page 209: #2-4, 6, 7, 10, 11, 13, 15, 18, 19, 21, 23, 24, 28, 30, 34a-c, 37, 39, 41-43, 45, 46, 48, 49, 52, 71





**Across**

**Down**

- |  |   |
|--|---|
| <p>2. bonds where more than one pair of electrons are shared between two atoms.</p> <p>5. the forces created by the unequal distribution of charge result in varying degrees of attraction between molecules</p> <p>7. elements under the staircase division in the periodic table</p> <p>8. the ability to carry an electrical current when electrons or ions are free to move</p> <p>9. an example of a strong intermolecular force</p> <p>10. electrons in covalent bonds are</p> <p>13. bonding between same nonmetals</p> <p>15. electrons in ionic compounds are</p> <p>19. 4 electrons are shared</p> <p>20. bonding between different nonmetal</p> <p>22. even on all sides</p> <p>24. this indicates how strongly an atom of an element attracts electrons in a chemical bond</p> | <p>1. ions that have both ionic and covalent bonding</p> <p>3. the energy required to remove an electron from a ground-state atom or ion.</p> <p>4. the molecular polarity can be determined by the _____ of charge</p> <p>6. 6 electrons are shared</p> <p>7. bonds formed when valence electrons are mobile</p> <p>9. the ability to resist physical changes</p> <p>11. a gas that has stable valence configurations and tends not to bond.</p> <p>12. elements above the staircase division in the periodic table.</p> <p>14. the ability of solids to be hammered into thin sheets</p> <p>16. when a bond is formed, energy is</p> <p>17. when a bond is broken, energy is</p> <p>18. molecular compound</p> <p>21. a bond between appositively charged ions</p> <p>23. diagrams representing the valence electrons</p> |
|--|---|